

# Chemistry Notes for class 12 Chapter 11

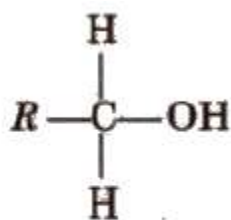
## Alcohols, Phenols and Ethers

### Alcohols and Phenols

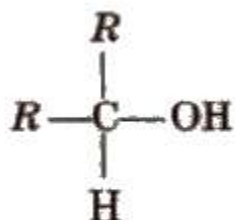
Alcohols and phenols are formed when a hydrogen atom in hydrocarbon, aliphatic and aromatic respectively, is replaced by hydroxyl group (-OR group).

### Classification of Alcohols and Phenols

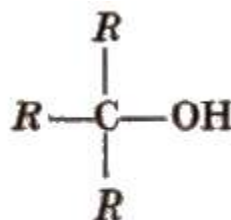
In alcohols, -OR group is attached to  $Sp^3$  hybridised carbon. These alcohols are usually classified as primary, secondary and tertiary alcohols.



primary ( $1^\circ$ )



secondary ( $2^\circ$ )



tertiary ( $3^\circ$ )

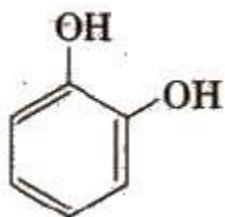
Alcohols may be

- (i) monohydric-containing one - OR group,
- (ii) dihydric-containing two - OR groups and
- (iii) polyhydric-containing three or more -OR groups.

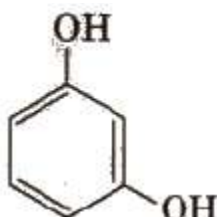
In phenols, -OR group is attached to  $Sp^2$  hybridised carbon. These may also be monohydric, dihydric, etc. The dihydric phenol further may be ortho, meta or para derivative.



phenol



catechol



resorcinol



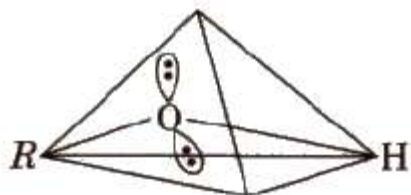
quinol or hydroquinone

In allylic alcohols, -OH group is attached to  $sp^3$  hybridised carbon but next to  $C=C$  bond.

e.g.,  $\text{CH}_2 = \text{CH} - \text{CH}_2\text{OH}$ , Benzylic alcohol ( $\text{C}_6\text{H}_5\text{CH}_2\text{OH}$ )

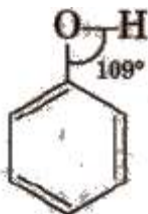
## Structure of Alcohols and Phenols

The oxygen atom of alcohols is  $\text{Sp}^3$  hybridised and they have tetrahedral position of hybrid atomic orbitals.



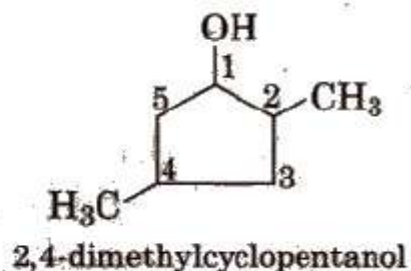
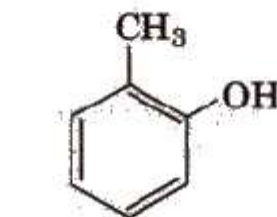
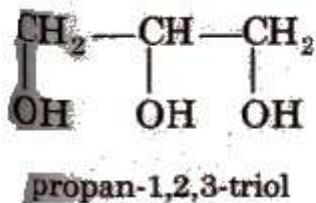
The value of  $\text{LROH}$  bond angle depends upon the R group. For methyl alcohol, it is ( $\angle \text{C} - \text{O} - \text{H}$ )  $108.9^\circ$  due to repulsion of lone pairs.

In phenols, the  $-\text{OH}$  group is attached to  $\text{Sp}^2$  hybridised carbon and thus, the  $\text{C} - \text{O}$  bond acquires a partial double bond character.



## Nomenclature of Alcohols and Phenol

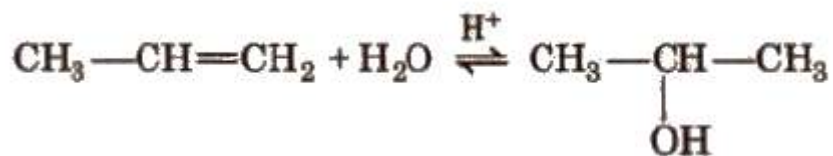
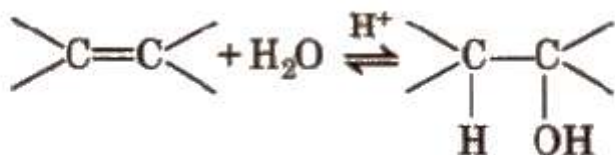
In IUPAC, system, alcohol or alkanols are named by replacing the last word 'e' of the corresponding alkane by 'ol'. e.g.,



## Preparation of Alcohols

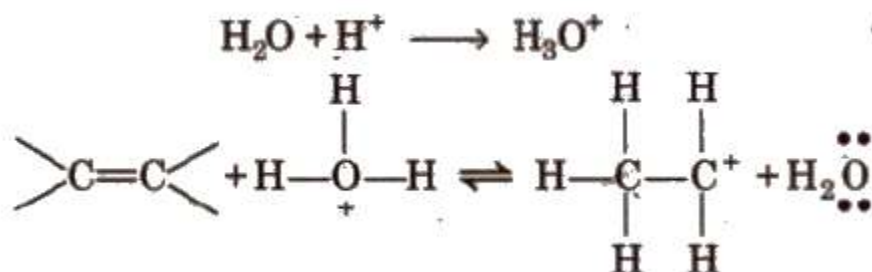
### (i) From alkenes

(a) By acid catalysed hydration in accordance with Markownikoff's rule.

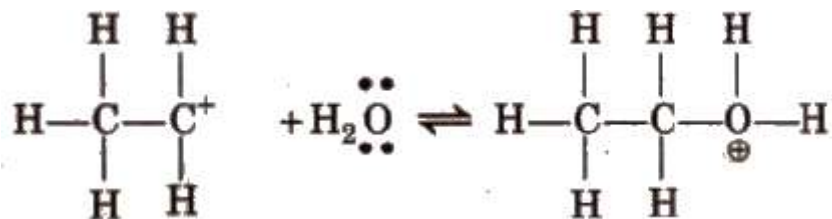


## Mechanism

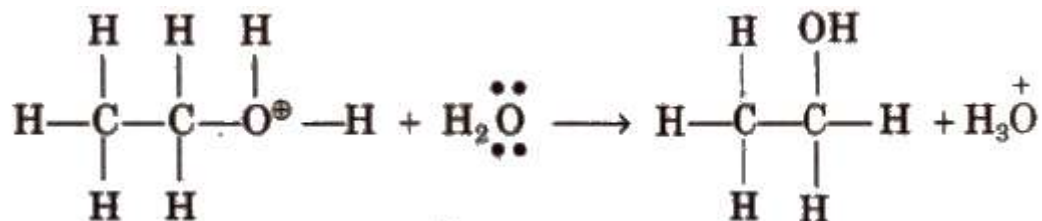
**Step I** Protonation of alkene by attack of  $\text{H}_3\text{O}^+$



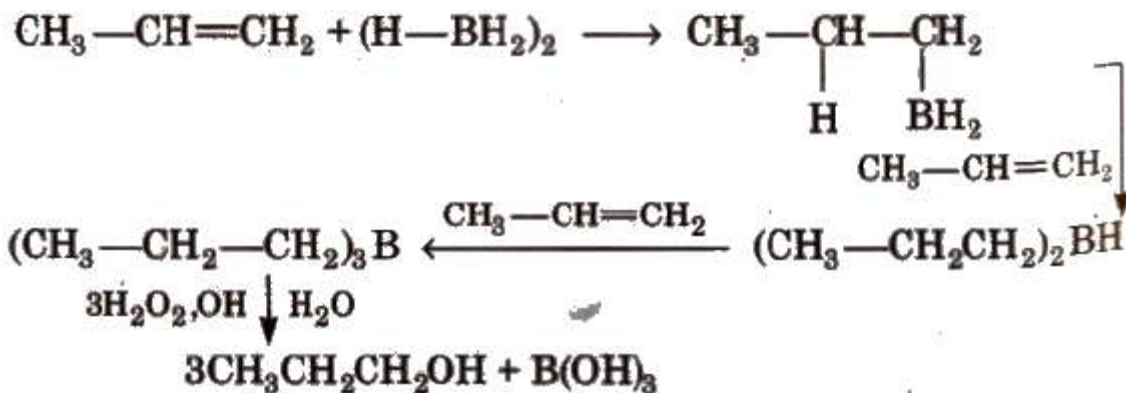
**Step II** Nucleophilic attack



**Step III** Deprotonation to form an alcohol

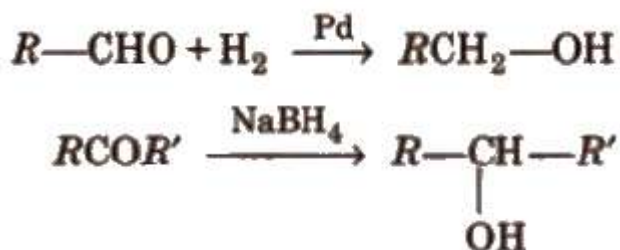


(b) By hydroboration-oxidation



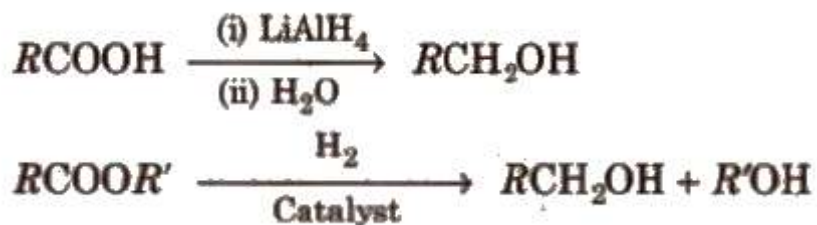
(ii) **From carbonyl compounds**

(a) By reduction of aldehydes and ketones

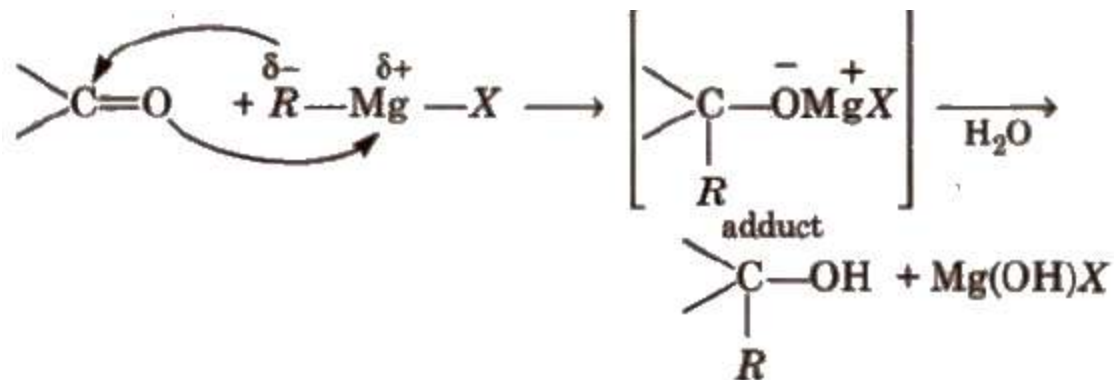


Aldehydes yield primary alcohols whereas ketones give secondary alcohols, when subjected to reduction.

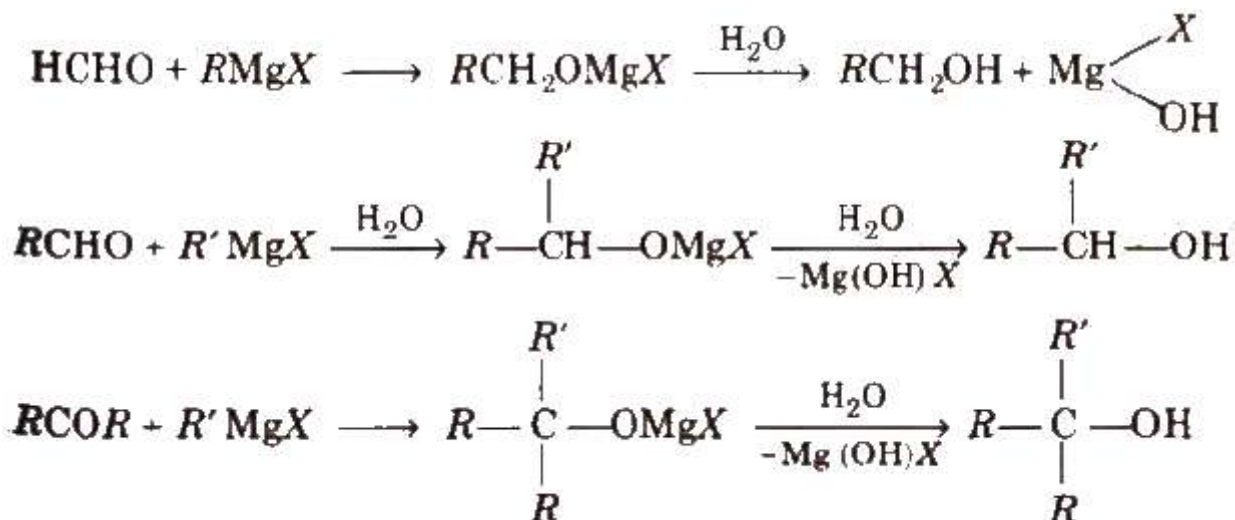
(b) By reduction of carboxylic acids and ester



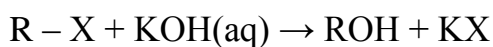
Reduction of aldehyde, ketones and esters with No Alcohol is called Bouveault-blanc reduction.



The reaction produces a primary alcohol with methanol, a secondary alcohol with aldehydes (except methanal) and tertiary alcohol with ketones



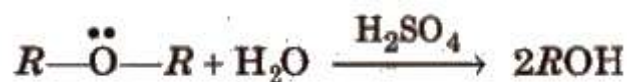
#### (iv) Hydrolysis of alkyl halides



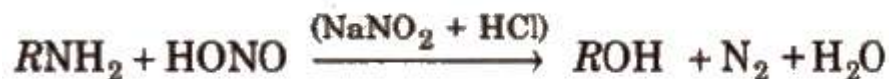
To avoid dehydrohalogenation of RX, mild alkalies like moist

Ease of hydrolysis of alkyl halides  $\text{RI} > \text{R}-\text{Br} > \text{RCI} >$  and  $\text{t} > \text{s} > \text{p}$  alkyl halides.

#### (v) Hydrolysis of ethers

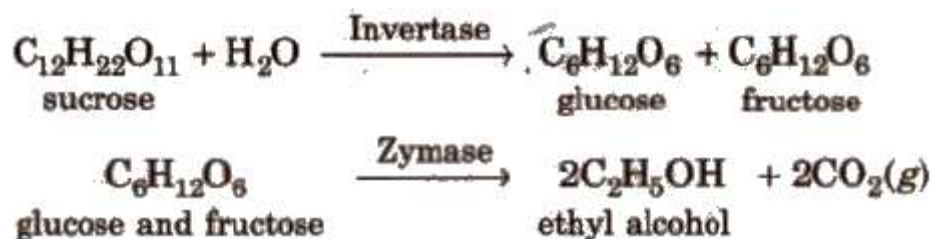


(vi) From primary amines By treatment with nitrous acid.



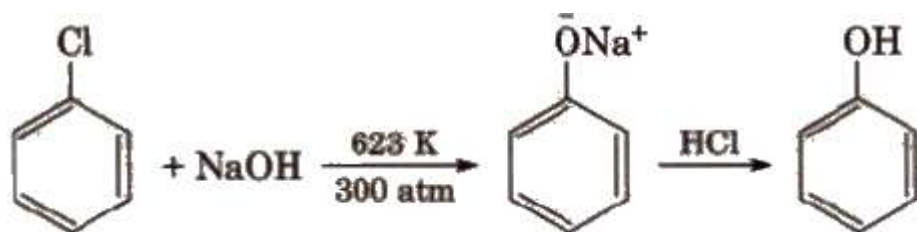
Methylamine does not give methyl alcohol when treated with  $\text{HNO}_2$ . It gives  $\text{CH}_3\text{OCH}_3$  and  $\text{CH}_3\text{ONO}$ .

(vii) **By alcoholic fermentation**

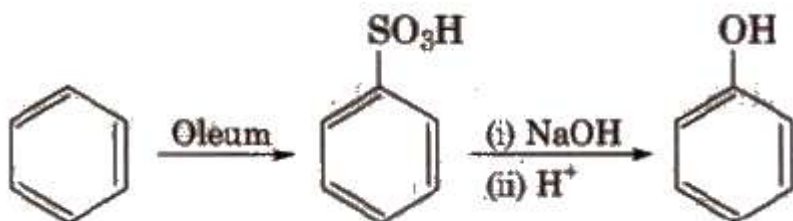


**Preparation of Phenols**

(i) **From haloarenes**



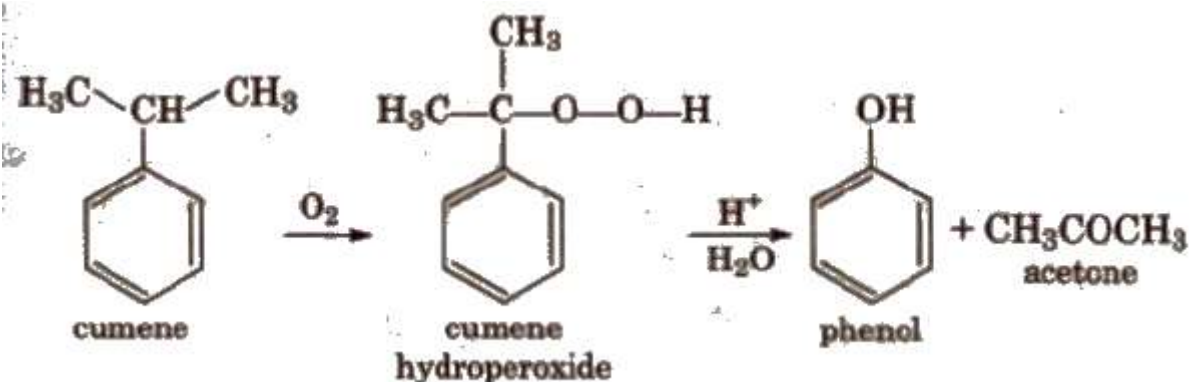
(ii) **From benzene sulphonic acid**



(iii) **From diazonium salts**

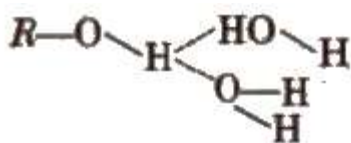


(iv) **From cumene**



## Physical Properties of Alcohols

1. Lower alcohols are colourless liquids, members from  $\text{C}_5 - \text{C}_{11}$  are oily liquids and higher members are waxy solids.
2. The hydroxyl groups in alcohols can form H-bonds with water, so alcohols are miscible with water. The solubility decreases with increase in molecular mass.



3. Boiling points of alkanes are higher than expected because of the presence of intermolecular hydrogen bonding in the polar molecules.

[The boiling point decreases in the order  $1^\circ > 2^\circ > 3^\circ$  as the van der Waals' forces of attraction decreases]

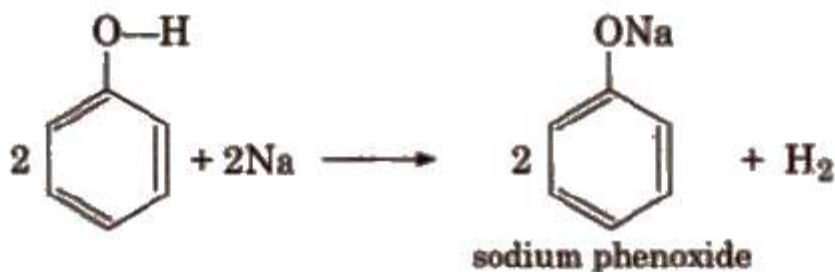
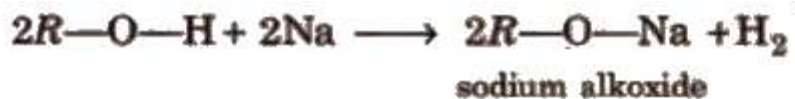
## Physical Properties of Phenols

1. These are colourless liquids or crystalline solids but become coloured due to slow oxidation with air.
2. Phenol is also called carbolic acid.
3. Because of the presence of polar -OH bond, phenols form intermolecular H-bonding with other phenol molecules and with water.

## Chemical Reactions of Alcohols and Phenols

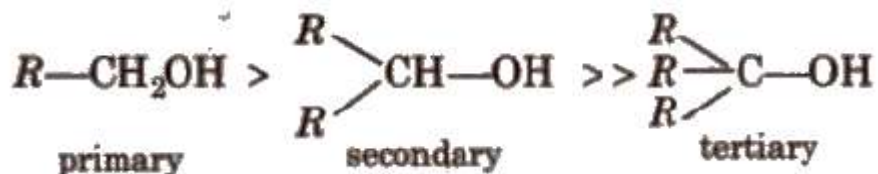
### (i) Reactions involving cleavage of O - H Bond

- (a) Acidity of alcohols and phenols



Alcohols are weaker acids than water due to +I group present in alcohols, which decreases the polarity of -O-H bond.

Acid strength of alcohols



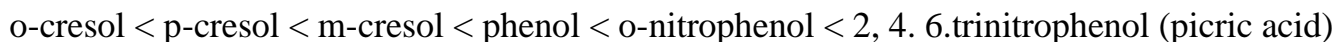
Electron releasing group increases electron density on oxygen to decrease the polarity of -OH bond.

Order of acidity is



Phenol is more acidic than alcohols due to stabilisation of phenoxide ion through resonance. Presence of electron withdrawing group increases the acidity of phenol by stabilising phenoxide ion while presence of electron releasing group decreases the acidity of phenol by destabilising phenoxide ion.

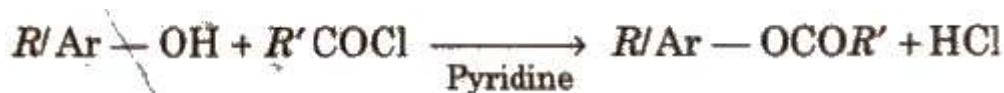
Thus, increasing acidic strength is



Higher  $K_a$  and lower  $pK_a$  value corresponds to the stronger acid.

(b) **Esterification**



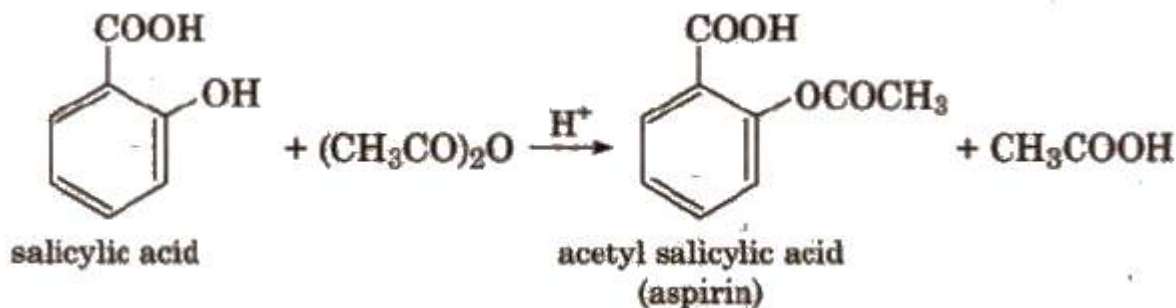


The reaction with R'COOH and (R' CO)<sub>2</sub>O is reversible, so cone, H<sub>2</sub>SO<sub>4</sub> is used to remove water.

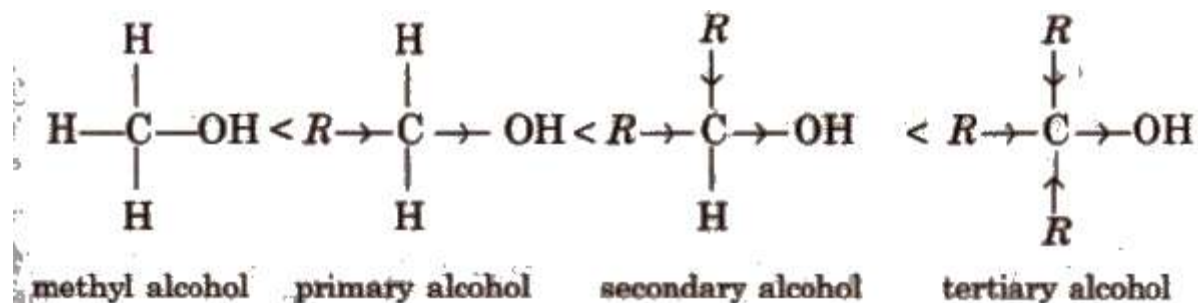
The reaction with R' COCl is carried out in the presence of pyridine so as to neutralise HCl which is formed during the reaction.

The introduction of acetyl (CH<sub>3</sub>CO-) group in phenols is known as acetylation.

Acetylation of salicylic acid produces aspirin.



(ii) **Reaction involving cleavage of C-O bond in alcohols** In these reactions, the reactivity order of different alcohols :



Alkyl group due to +I effect increases the electron density on the carbon and oxygen atom of C-OH bond. As a result, the bond cleavage becomes easy. Greater the number of alkyl groups present, more will be the reactivity of alcohol. Thus, the relative order of reactivity of the alcohols is justified.

(a) **Reaction with halogen acids** Alcohols can be converted into haloalkanes by the action of halogen acids.



For a given alcohol order of reactivity of HX is



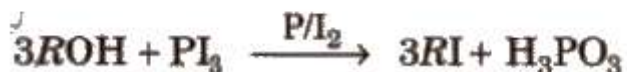
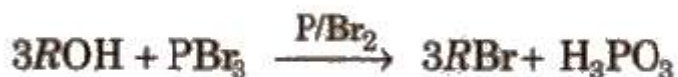
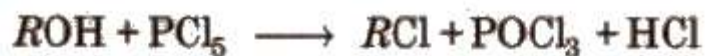
For a given halogen acid order of reactivity of alcohols

Tertiary > Secondary > Primary

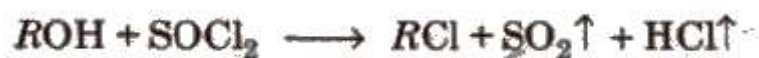
**Lucas test**

Primary alcohols	Secondary alcohols	Tertiary alcohols
$RCH_2OH \xrightarrow[\text{Anhy. ZnCl}_2]{\text{Conc HCl}}$	$R_2CH-OH \xrightarrow[\text{Anhy. ZnCl}_2]{\text{Conc HCl}}$	$R_3C-OH \xrightarrow[\text{Anhy. ZnCl}_2]{\text{Conc HCl}}$
No reaction and hence, no white cloudiness or turbidity at room temperature.	$R_2CHCl$ White cloudiness or turbidity appears with in about 5 minutes.	$R_3CCl$ White cloudiness or turbidity appears immediately.

(b) **Reaction with phosphorus halides**

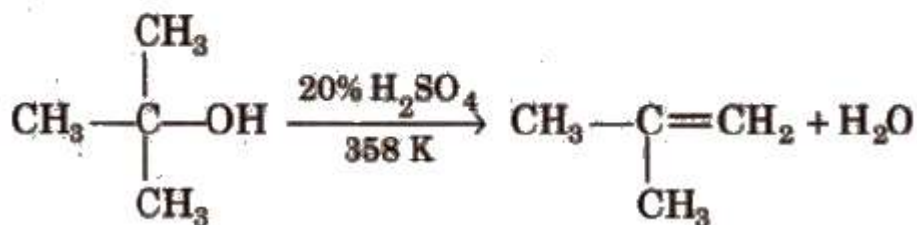
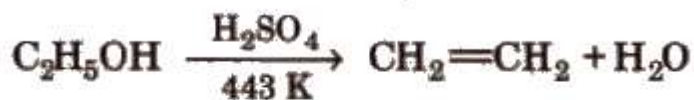
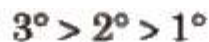


(c) **Reaction with thionyl chloride**



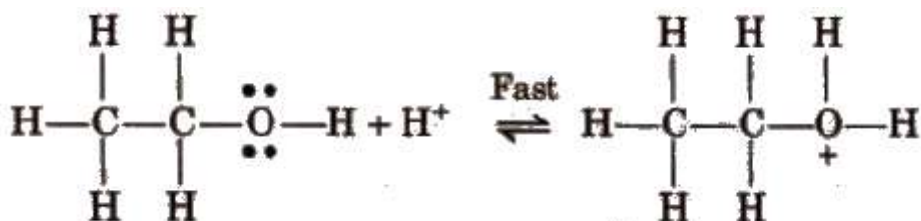
d) **Dehydration of alcohols** It requires acid catalyst and the reaction proceeds via intermediate carbonium ion. Acidic catalyst converts hydroxyl group into a good leaving group.

Since, the rate determining step is the formation of carbocation, the ease of dehydration is

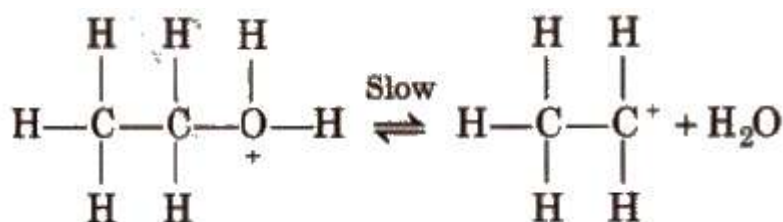


## Mechanism

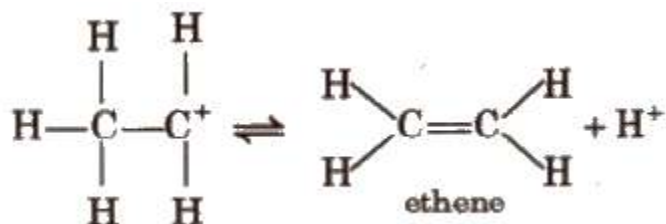
### Step I Formation of protonated alcohol



### Step II Formation of carbocation



### Step III Formation of ethene by elimination of a proton

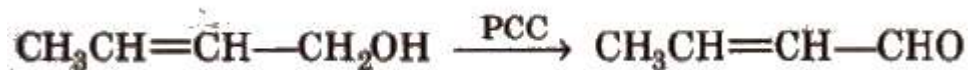


In dehydration reaction, highly substituted alkene is the major product and if the major product is capable of showing cis-trans isomerism, trans-product is the major product. (Saytzeff's rule).

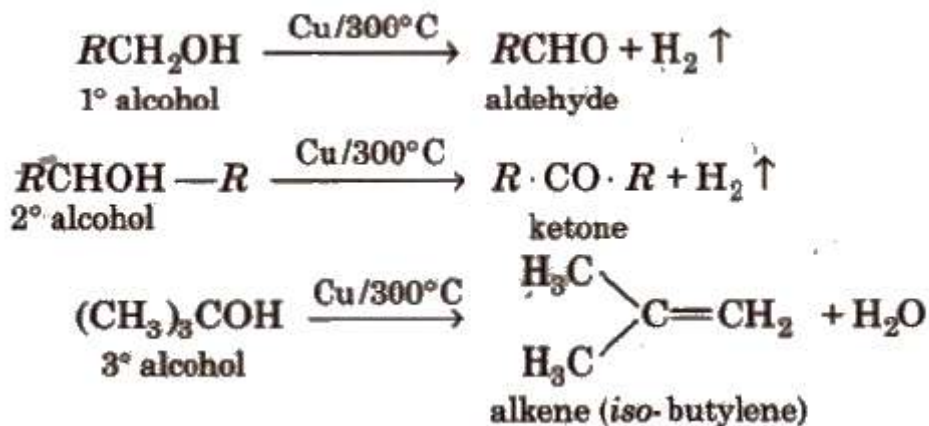
(iii) Oxidation reactions Oxidising reagents used for the oxidation of alcohols are neutral, acidic or alkaline  $\text{KMnO}_4$  and acidified  $\text{K}_2\text{Cr}_2\text{O}_7$ .

Primary alcohols	Secondary alcohols	Tertiary alcohols
$\text{CH}_3\text{CH}_2\text{OH}$	$\text{CH}_3\text{CHOH}-\text{CH}_3$	$(\text{CH}_3)_3\text{C}-\text{OH}$
$\downarrow[\text{O}]$	$\downarrow[\text{O}]$	$\downarrow[\text{O}]$
$\text{CH}_3\text{CHO}$	$\text{CH}_3\text{COCH}_3$	$\text{CH}_3\text{COCH}_3 + \text{CO}_2 + \text{H}_2\text{O}$
$\downarrow[\text{O}]$	$\downarrow[\text{O}]$	$\downarrow[\text{O}]$
$\text{CH}_3\text{COOH}$	$\text{CH}_3\text{COOH} + \text{CO}_2 + \text{H}_2\text{O}$	$\text{CH}_3\text{COOH} + \text{CO}_2 + \text{H}_2\text{O}$

A common reagent that selectively oxidises a primary alcohol to an aldehyde (and no further) is pyridinium chlorochromate (PCC).



#### (iv) Dehydrogenation

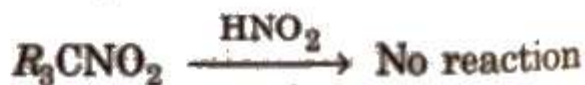
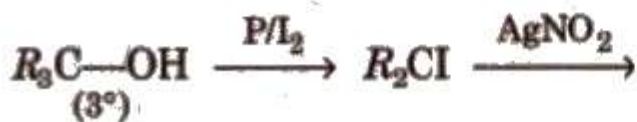
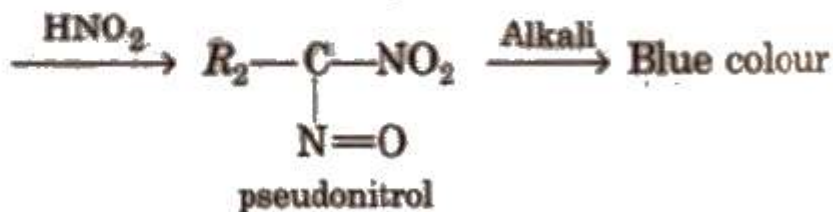
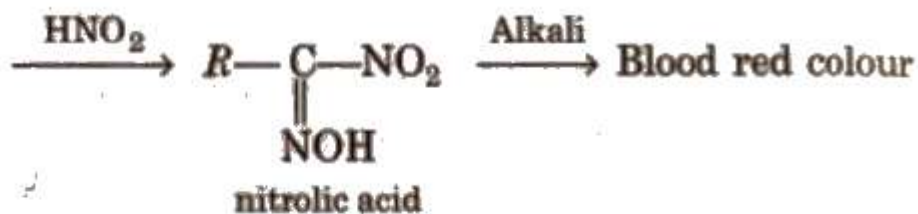


#### Distinction among 1°, 2° and 3° Alcohols

1°, 2° and 3° alcohols are distinguished by Lucas test, oxidation and reduced copper.

Victor Meyer's test is also used to distinguish them.

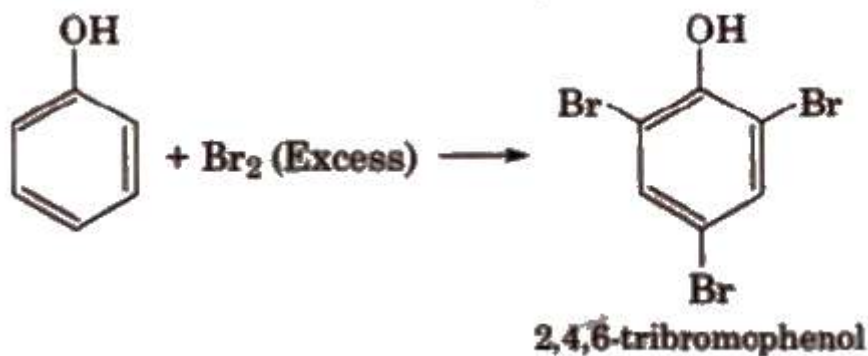
In this test, primary (1°) alcohols give red colour, secondary (2°) alcohols give blue colour and tertiary (3°) alcohols give no colouration.



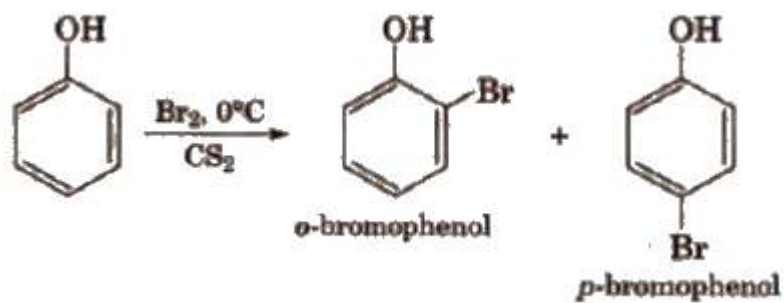
## Reactions of Phenols

(i) **Electrophilic substitution reactions** The -OH group attached to the benzene ring activates it towards electrophilic substitution at ortho and para positions .

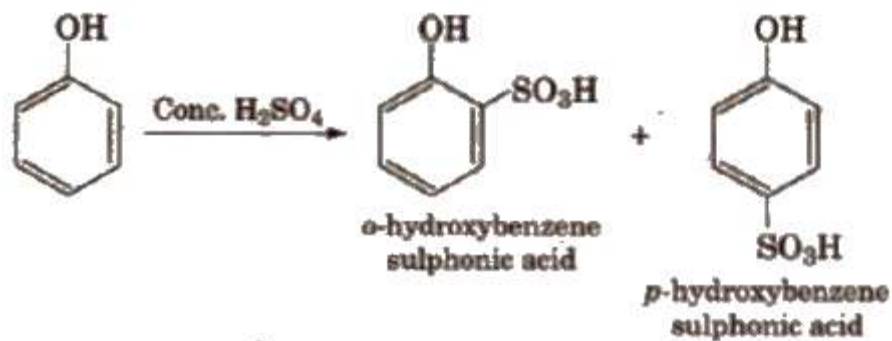
(a) **Halogenation**



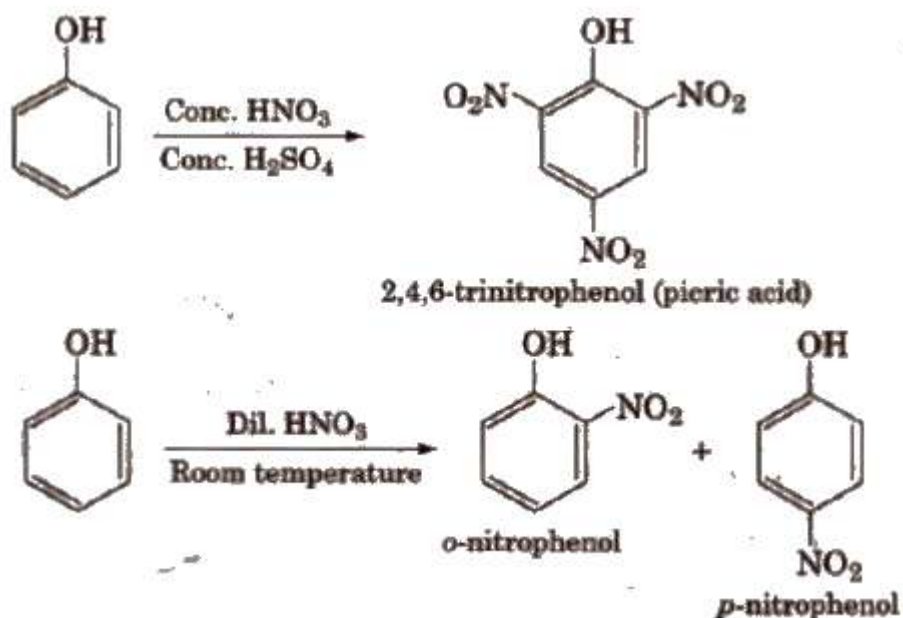
With calculated amount of Br<sub>2</sub> in CS<sub>2</sub> or CHCl<sub>3</sub> it gives ortho and para product.



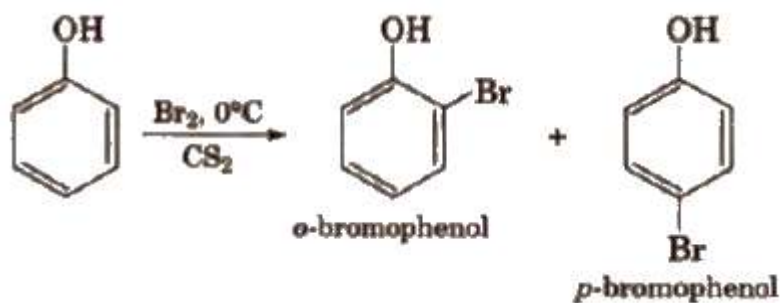
### (b) Sulphonation



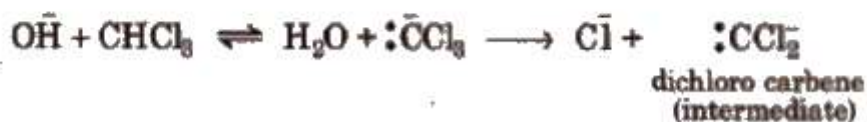
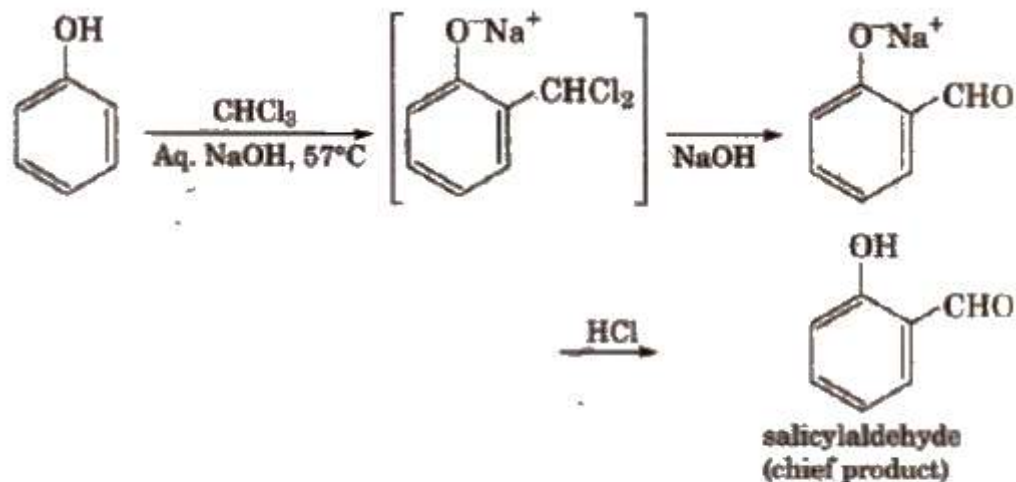
### (c) Nitration



The ortho and para isomers can be separated by steam distillation. This is because *o*-nitrophenol is steam volatile due to intramolecular hydrogen bonding while *p*-nitrophenol is less volatile due to intermolecular hydrogen bonding which causes the association of molecules.

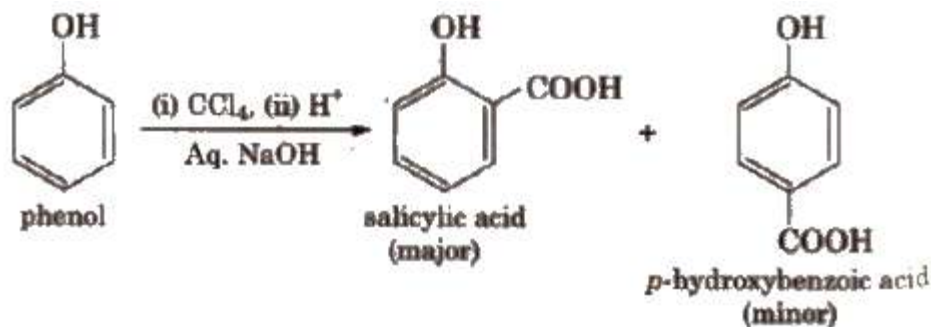


(d) **Reimer-Tiemann reaction**

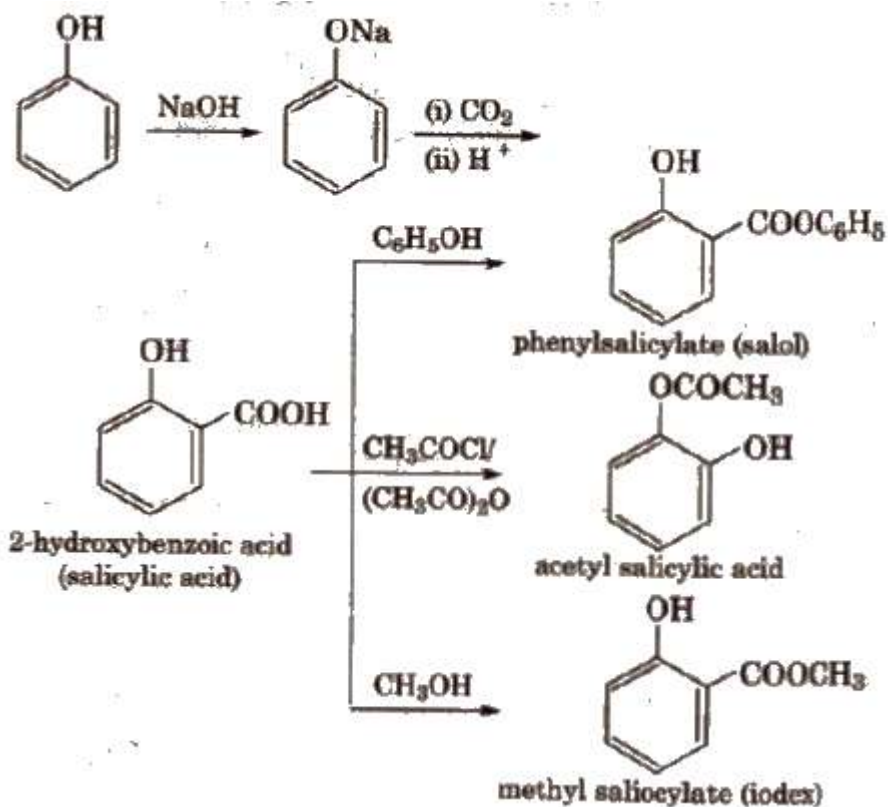


This reaction is an electrophilic substitution reaction and electrophile is dichlorocarbene.

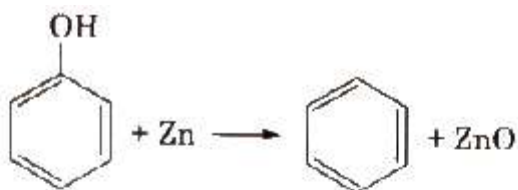
Similarly with carbon tetrachloride and alkali, *c*- and *p*-hydroxybenzoic acid are obtained



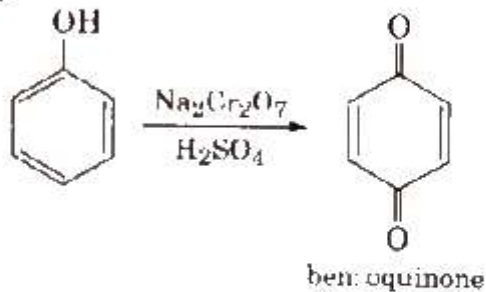
(ii) **Kolbe's reaction**



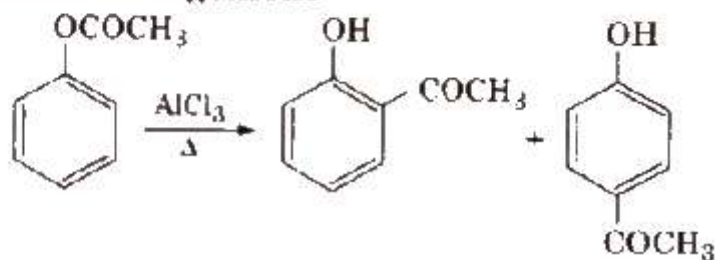
(iii) Reaction with zinc dust



(iv) Oxidation



(v) Fries rearrangement





(a) **Rectified spirit** It contains 95% ethyl alcohol and 45% water. It is an azeotrope (constant boiling mixture) and boils at 74°.

(b) **Absolute alcohol** Alcohol containing no water, i.e; 100% C<sub>2</sub>H<sub>5</sub>OH is known as absolute alcohol. It is prepared as follows.

(i) Quick lime process

(ii) Azeotropic method

(c) **Methylated spirit** The rectified spirit rendered poisonous by addition of 4-5% methyl alcohol, traces of pyridine and some copper sulphate and is known as methylated spirit or denatured alcohol.

(d) **Power alcohol** Alcohol mixed with petrol or fuel and used In internal combustion engines Is known as power alcohol.

(e) **Wood spirit** Methyl alcohol (CH<sub>3</sub>OH) is also called wood spirit. It is obtained by destructive distillation of wood. Pyroligneous add, the product of destructive distillation of wood, contains acetic acid (10%), methyl alcohol (25%) and acetone (05%). Drinking of methanol causes blindness.

(f) **Grain alcohol** Ethyl alcohol C<sub>2</sub>H<sub>5</sub>OH is also called grain alcohol. It is used In the preparation of various beverages containing different percentages.

## Dihydric Alcohols

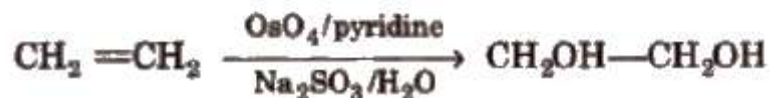
These are generally called glycols because of their sweet taste. Ethylene glycol (CH<sub>2</sub>OH – CH<sub>2</sub>OH) is the first and most important member of dihydric alcohol series.

## Methods of Preparation

(i) **From ethylene**



(1% alkaline KMnO<sub>4</sub> is called Baeyer's reagent)



(ii) **By reduction of glyoxal**

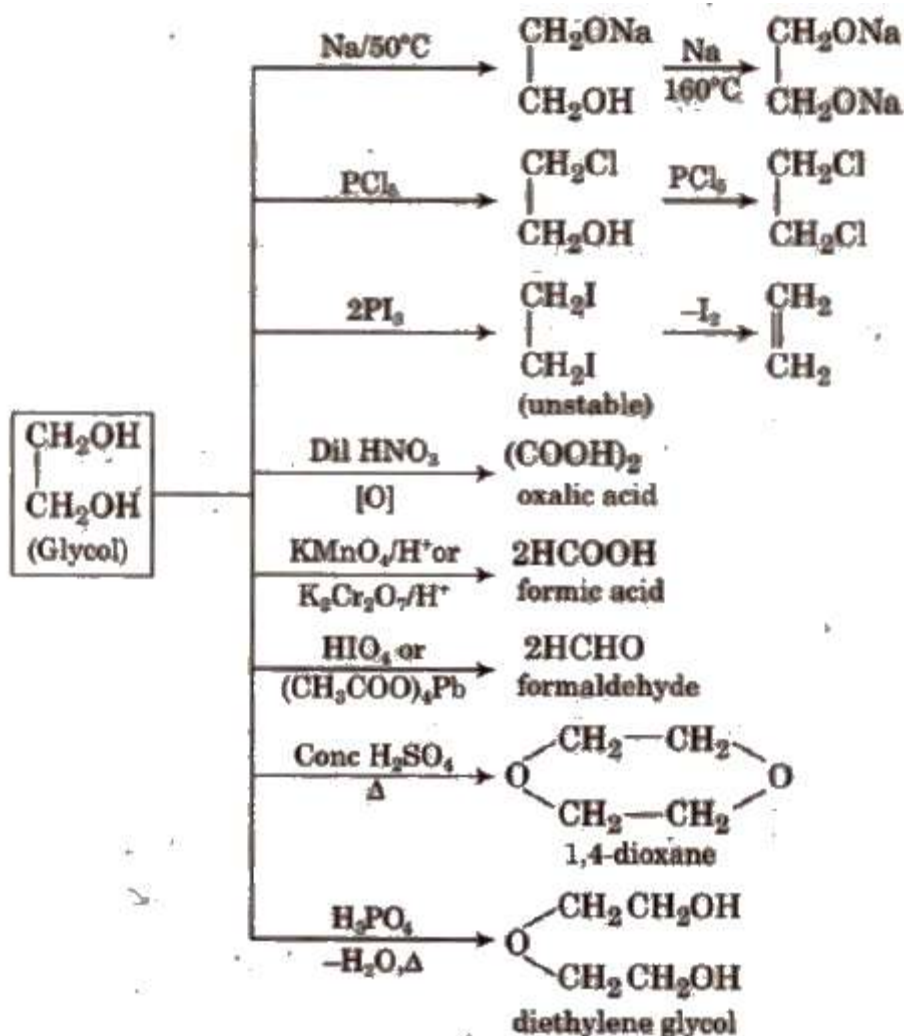


## Physical Properties

1. It is a colourless, syrupy liquid with sweet taste.
2. Because of its tendency of formation of H-bonds, it is miscible with H<sub>2</sub>O and ethanol but not with ether.

## Chemical Properties

It gives all the general reactions of -OH group.



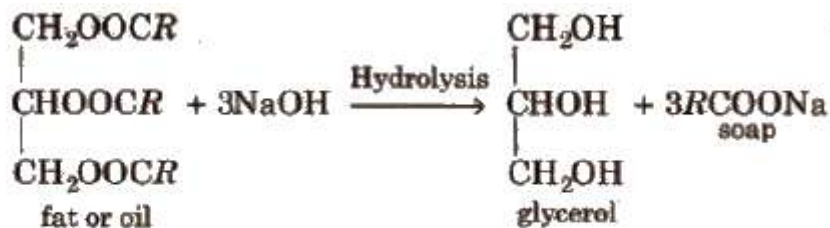
The per-iodic acid cleavage of 1,2-glycols is sometimes called Malaprade reaction.

## Trihydric Alcohols

Glycerol or glycerine, CH<sub>2</sub>OH – CH(OH) – CH<sub>2</sub>OH is the first member of this group. Its IUPAC name is propane-1,2,3-triol.

## Method of Preparation

It is obtained as a by product in saponification reaction.



(where,  $R = \text{C}_{17}\text{H}_{35}$  or  $\text{C}_{15}\text{H}_{31}$  or  $\text{C}_{17}\text{H}_{33}$ )

## Physical Properties

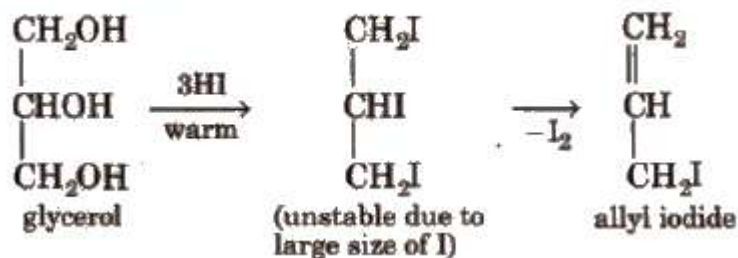
1. It is a colourless, odourless, viscous and hygroscopic liquid.
2. It is sweet in taste and steam volatile.
3. It is soluble in water but insoluble in ether.
4. Due to excessive H-bonding, it is highly viscous and has high boiling point.

## Chemical Properties

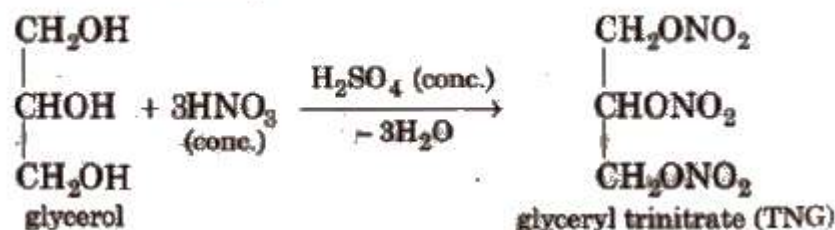
It gives all the general reactions given by -OR group but 2° OR is less reactive as compared to 1°.

Some of its specific reactions are :

### (i) Reaction with HI



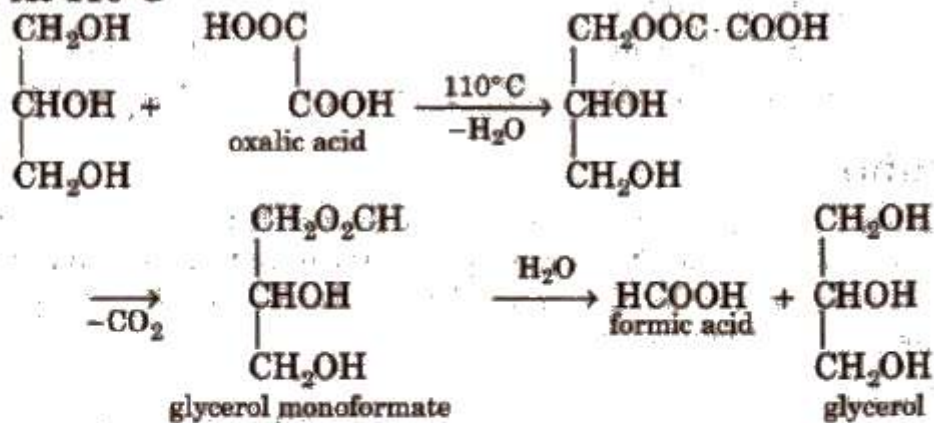
### (ii) Reaction with $\text{HNO}_3$



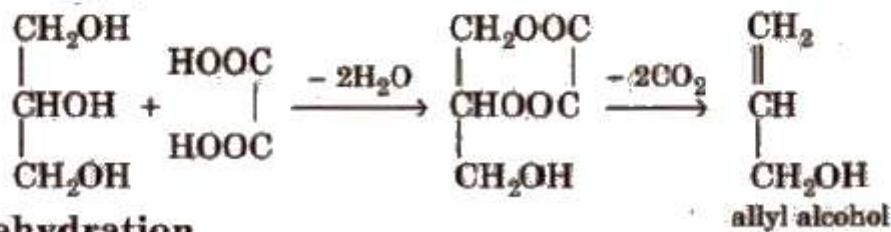
Glyceryl trinitrate or tri nitroglycerine, when adsorbed on Kieselguhr is known as dynamite. Mixture of TNG and cellulose trinitrate is called blasting gelatin.

(iii) Reaction with oxalic acid

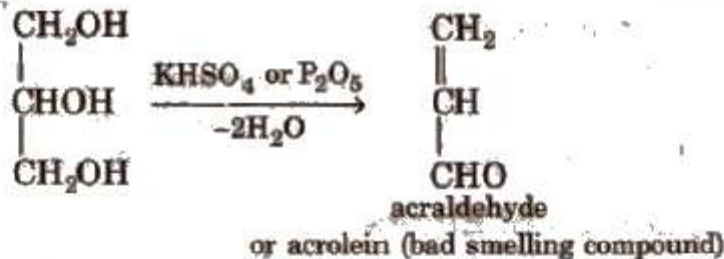
(a) At 110°C



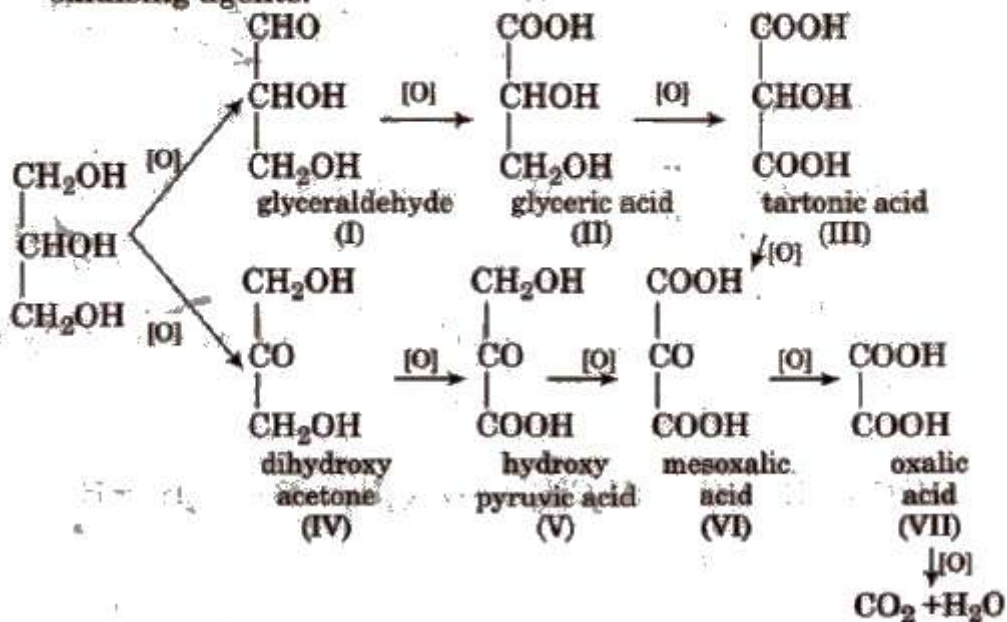
(b) At 260°C



(iv) Dehydration



(v) Oxidation Different products are obtained by different oxidising agents.



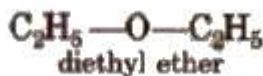
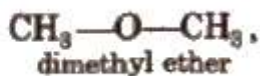
Cone  $\text{HNO}_3$  gives II; dil  $\text{HNO}_3$  gives II and III;  $\text{Bi}(\text{NO}_3)_3$  or  $\text{NaNO}_3$  gives VI; Fenton's reagent or  $\text{NaOBr}$  or  $\text{Br}_2$  water in  $\text{Na}_2\text{CO}_3$  gives a mixture of I and IV.

Solid  $\text{KMnO}_4$  oxidises glycerol to VII and  $\text{CO}_2$  and  $\text{H}_2\text{O}$ .

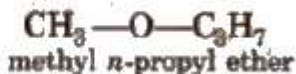
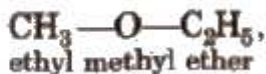
With  $\text{HIO}_4$  (periodic acid), glycerol gives  $\text{HCOOH}$  and  $\text{HCHO}$ .

## Ethers

Ethers are the organic compounds in which two alkyl or aryl groups are attached to a divalent oxygen, known as ethereal oxygen. These are represented by the general formula  $\text{R}-\text{O}-\text{R}'$  where R may be alkyl or aryl groups. e.g.,



### Mixed ethers



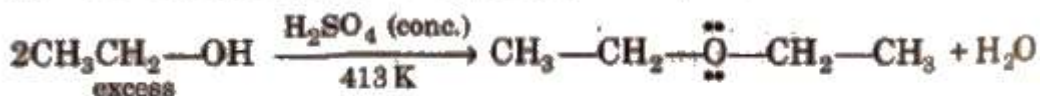
These are the functional isomers of alcohols. These also exhibit chain isomerism and metamerism.

## Nomenclature of Ethers

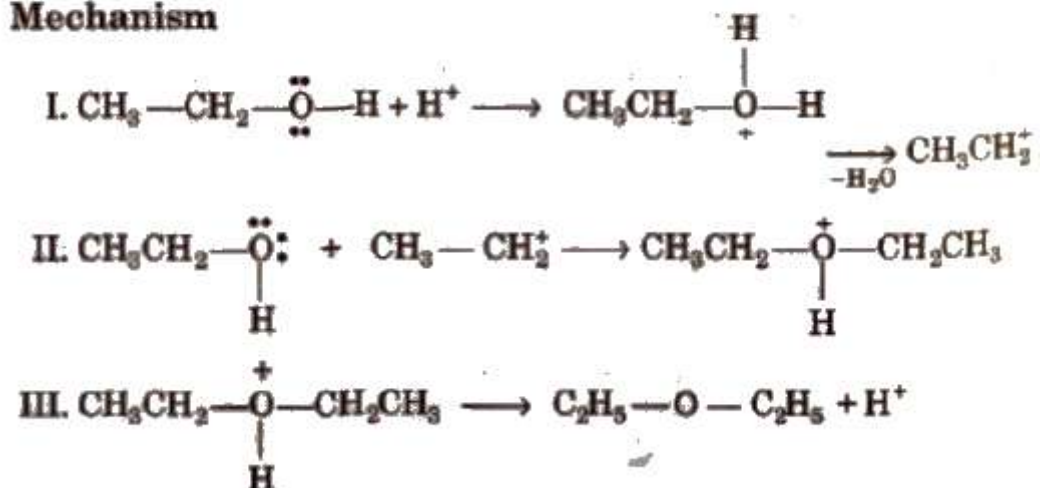
In the IUPAC system, ethers are regarded as 'alkoxy alkanes' in which the ethereal oxygen is taken along with smaller alkyl group while the bigger alkyl group is regarded as a part of the alkane.

## Preparation of Ethers

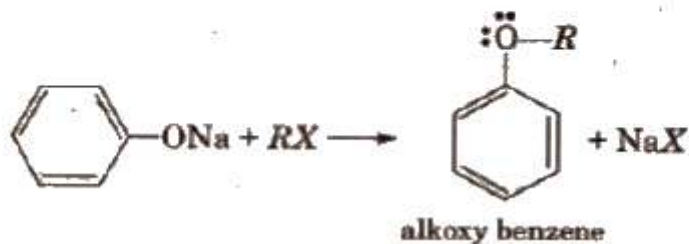
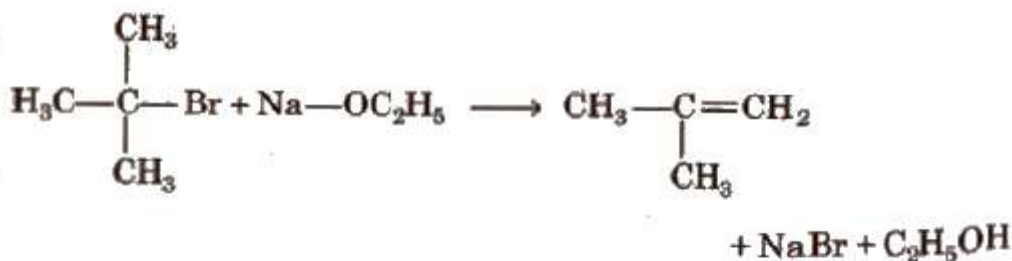
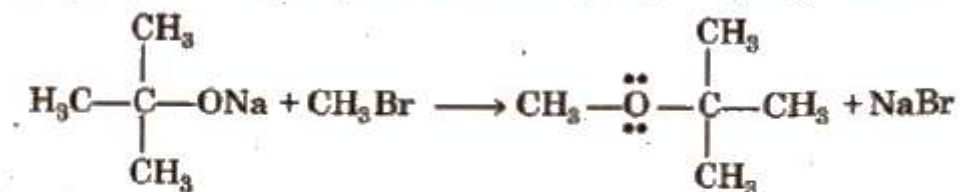
(i) **By dehydration of alcohols**



**Mechanism**



(ii) **Williamson's synthesis** Only primary alkyl halides when react with sodium alkoxide give ether while tertiary alkyl halides give alkene due to steric hindrance.

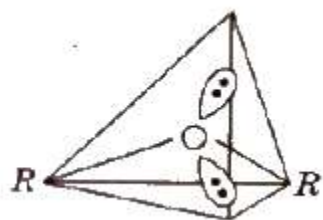


**Physical Properties of Ethers**

Ethers are polar but insoluble in  $\text{H}_2\text{O}$  and have low boiling point than alcohols of comparable molecular masses because ethers do not form hydrogen bonds with water.

## Structure of Ether

The hybridisation of O atom in ethers is  $sp^3$  (tetrahedral) and its shape is V-shape.

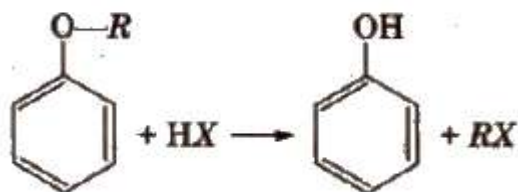


For dimethyl ether

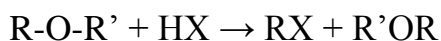


## Chemical Reactions of Ether

### (i) Reaction with HX



Ethers with two different alkyl groups are also cleaved in the same manner and results in the formation of a primary halide (or smaller and less complex alkyl halide) by  $S_N^2$  mechanism.



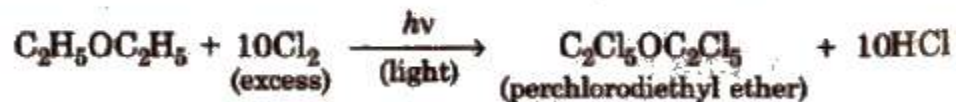
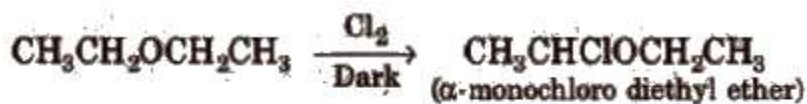
The order of reactivity of hydrogen halides is as follows



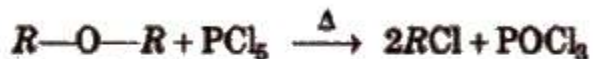
In ethers if one of the alkyl groups is a tertiary group, the halide formed is a tertiary halide by  $S_N^1$  mechanism.



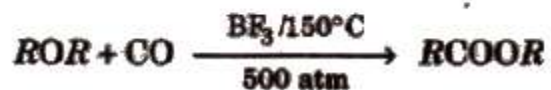
### (ii) Halogenation



(iii) Reaction with  $\text{PCl}_5$

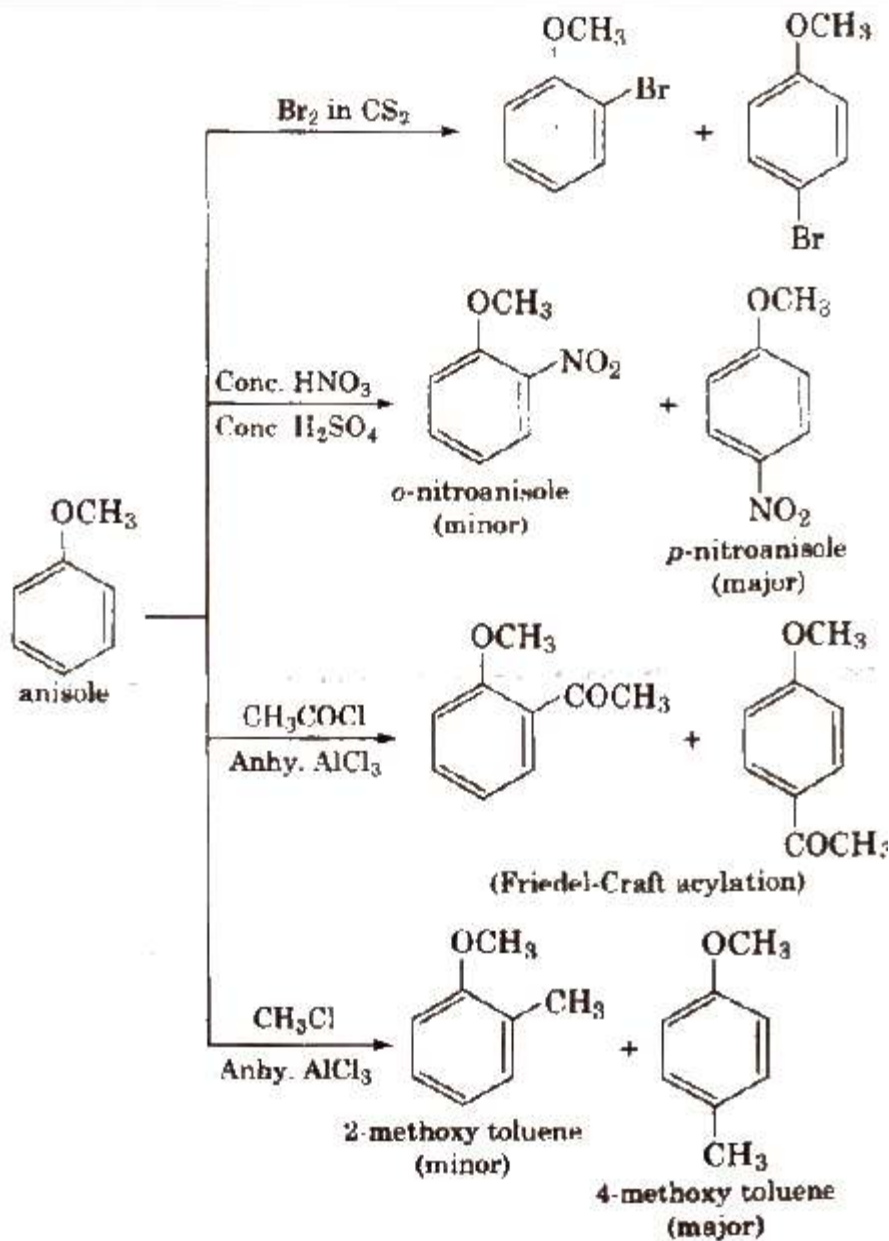


(iv) Reaction with CO



(v) Electrophilic substitution reactions In ethers, -OR is ortho, para directing group and activate the aromatic ring towards electrophilic substitution reaction.





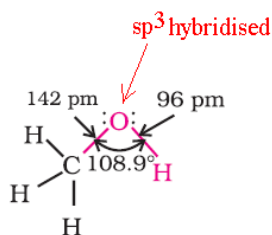
Ethyl phenyl ester  $\text{C}_6\text{H}_5\text{OC}_2\text{H}_5$  is also, known as phenetole.

### Uses of Ethers

1. Dimethyl ether is used as refrigerant and as a solvent at low temperature.
2. Diethyl Ether is used as an anaesthesia in surgery .

SUMMARY

• **Structure of alcohols:**

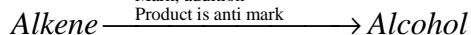


Methanol  
(Alcohol)

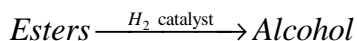
• **Preparation of alcohols:**

a) From alkene

Acid catalysed hydration  
( $H_2O, H^+$ )  
Mark, Addition  
Or  
Hydroboration - oxidation  
 $B_2H_6, H_2O_2 / OH^-$   
Mark, addition  
Product is anti mark

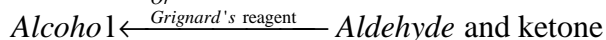


b) From esters

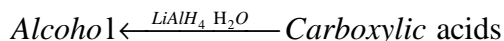


c) From aldehydes and ketones

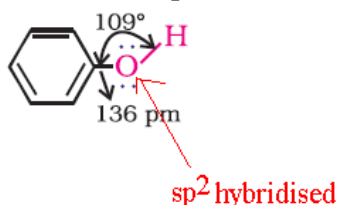
$H_2 / Pd$   
or  
 $NaBH_4$   
Or  
 $LiAlH_4$   
Or  
Grignard's reagent



d) From carboxylic acids



• **Structure of phenols:**



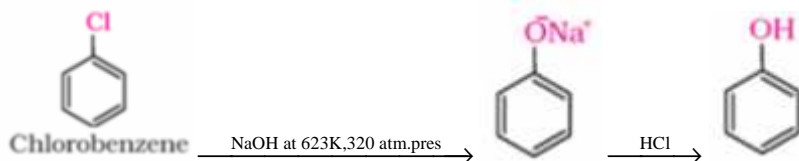
Phenol  
(Phenol)

• **Preparation of phenols:**

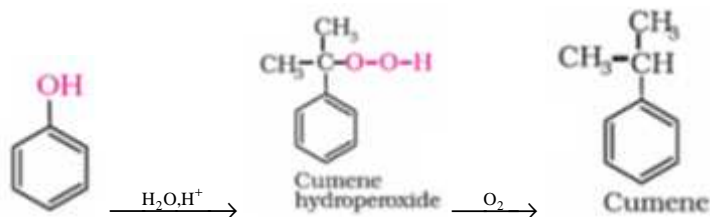
a) From benzene



b) From chlorobenzene



c) From cumene



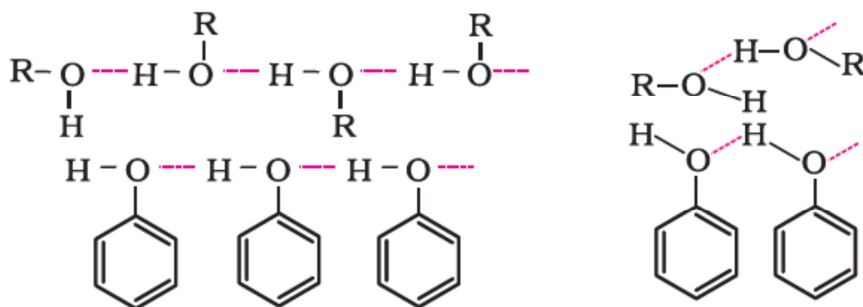
d) From aniline



- Physical properties of alcohols and phenols:

a) Boiling points

Boiling points of alcohols and phenols are higher in comparison to other classes of compounds. This is because the -OH group in alcohols and phenols is involved in intermolecular hydrogen bonding.



The boiling points of alcohols and phenols increase with increase in the number of carbon atoms. This is because of increase in van der Waals forces with increase in surface area.

In alcohols, the boiling points decrease with increase of branching in carbon chain. This is because of decrease in van der Waals forces with decrease in surface area.

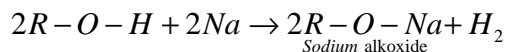
b) Solubility

Solubility of alcohols and phenols are soluble in water due to their ability to form hydrogen bonds with water molecules. The solubility of alcohols decreases with increase in size of alkyl/aryl (hydrophobic) groups.

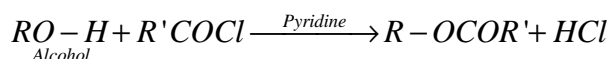
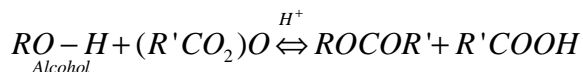
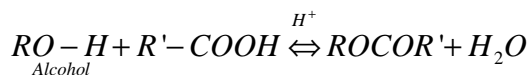
- **Chemical properties of alcohols:**

I. Reactions involving cleavage of O–H bond: Alcohols react as nucleophiles:

a) Reaction with metals

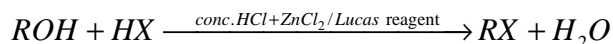


b) Esterification reaction

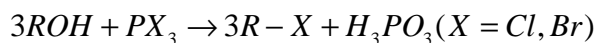


II. Reactions of alcohols involving cleavage of carbon – oxygen (C–O) bond:

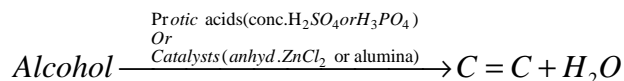
a) Reaction with hydrogen halides



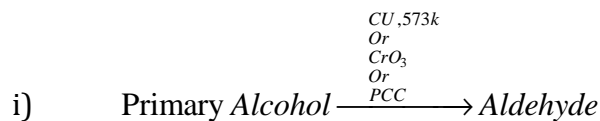
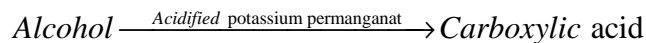
b) Reaction with phosphorus trihalides



c) Dehydration reaction



d) Oxidation reaction



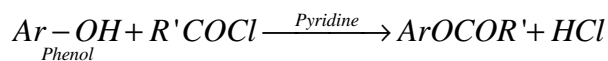
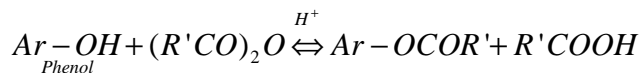
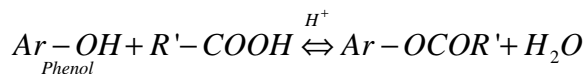
- **Chemical properties of phenols:**

I. Reactions involving cleavage of O–H bond: Alcohols react as nucleophiles:

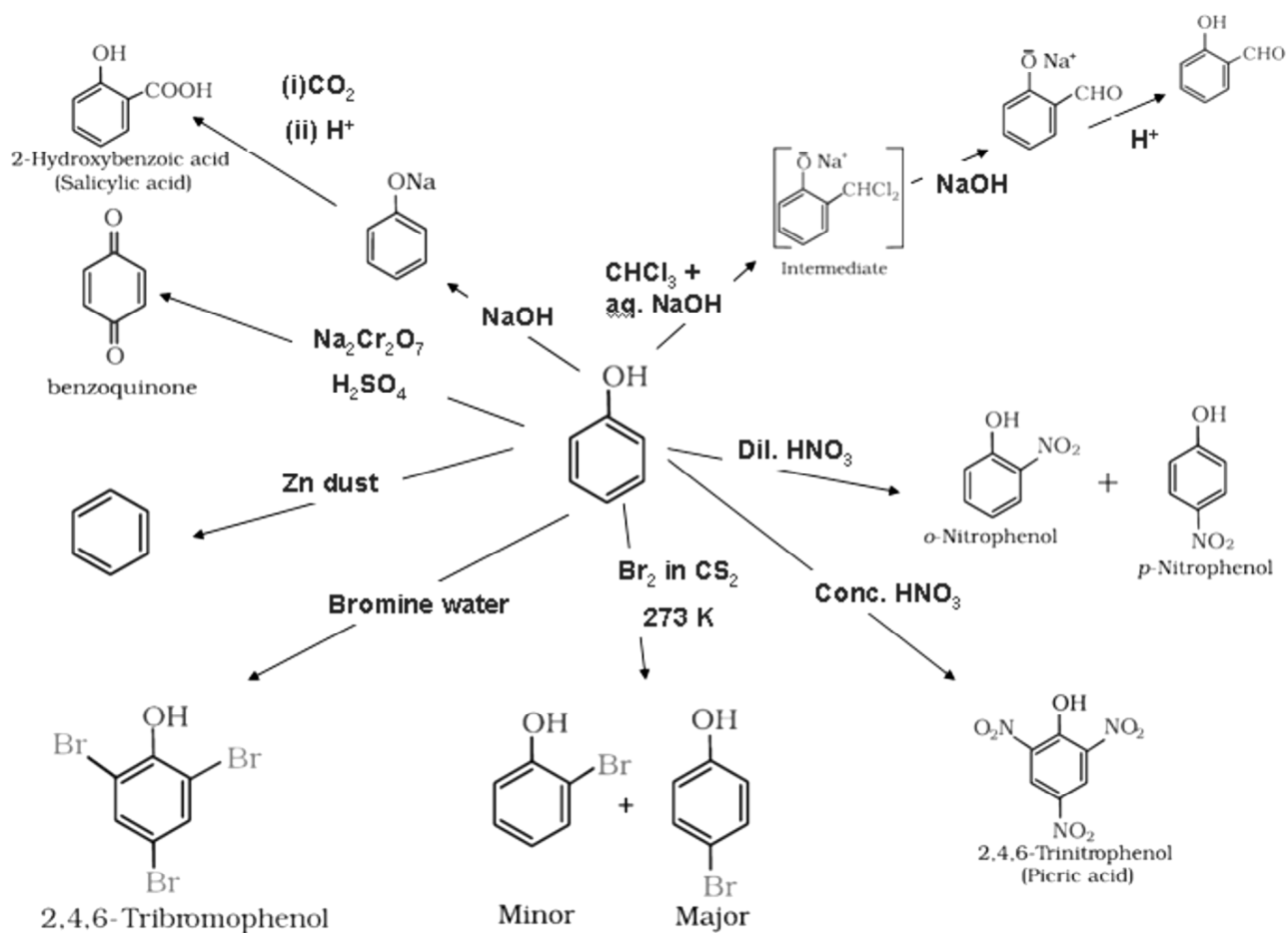
a) Reaction with metals



b) Esterification reaction



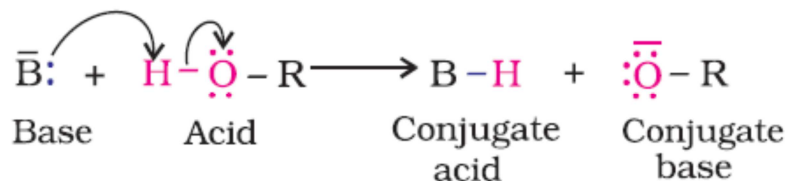
II. Other chemical reactions of phenols:



III. Acidic nature of phenol and alcohol:

a) Phenol >  $\text{H}_2\text{O}$  > Primary alcohol > Secondary alcohol > Tertiary alcohol

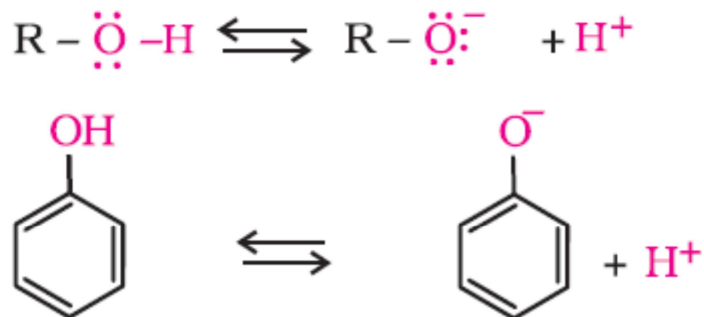
The acidic character of alcohols is due to the polar nature of O-H bond. Alkyl group is an electron-releasing group ( $-\text{CH}_3$ ,  $-\text{C}_2\text{H}_5$ ) or it has electron releasing inductive effect (+I effect). Due to +I effect of alkyl groups, the electron density on oxygen increases. This decreases the polarity of O-H bond. And hence the acid strength decreases.



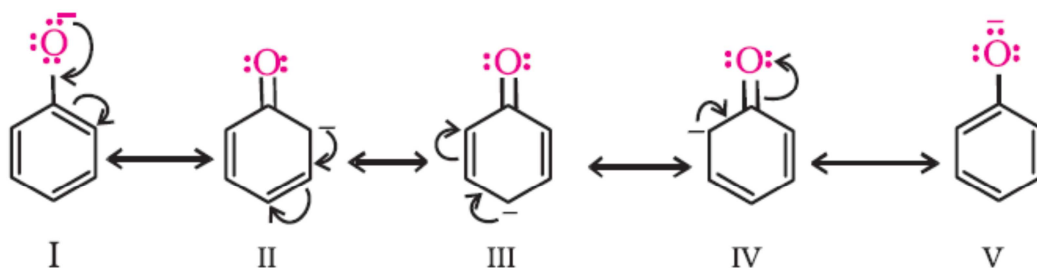
b) Phenol is more acidic than alcohol:

In phenol, the hydroxyl group is directly attached to the  $sp^2$  hybridised carbon of benzene ring which acts as an electron withdrawing group whereas in alcohols, the hydroxyl group is attached to the alkyl group which have electron releasing inductive effect. In phenol, the hydroxyl group is directly attached to the  $sp^2$  hybridised carbon of benzene ring whereas in alcohols, the hydroxyl group is attached to the  $sp^3$  hybridised carbon of the alkyl group. The  $sp^2$  hybridised carbon has higher electronegativity than  $sp^3$  hybridised carbon. Thus, the polarity of O-H bond of phenols is higher than those of alcohols. Hence, the ionisation of phenols is higher than that of alcohols.

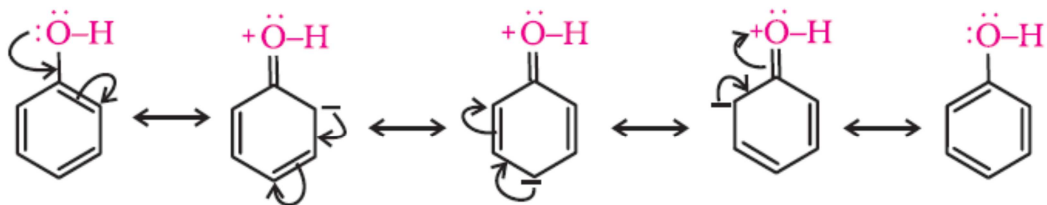
The ionisation of an alcohol and a phenol takes place as follows:



In alkoxide ion, the negative charge is localised on oxygen while in phenoxide ion, the charge is delocalised.



The delocalisation of negative charge makes phenoxide ion more stable and favours the ionisation of phenol. Although there is also charge delocalisation in phenol, its resonance structures have charge separation due to which the phenol molecule is less stable than phenoxide ion.

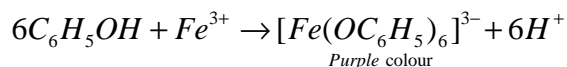


- c) In substituted phenols, the presence of electron withdrawing groups such as nitro group enhances the acidic strength of phenol. On the other hand, electron releasing groups, such as alkyl groups, in general, decreases the acid strength. It is because electron withdrawing groups lead to effective delocalisation of negative charge in phenoxide ion.

- **Differentiate between organic compounds:**

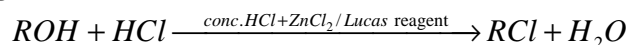
- a) Alcohols and phenols

Phenol on reaction with neutral  $\text{FeCl}_3$  gives purple colour whereas alcohols do not give purple colour.



- b) Primary, secondary and tertiary alcohols

Lucas reagent test:



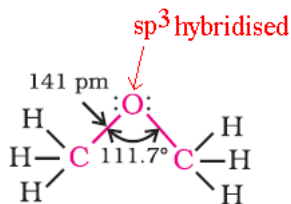
If it is a primary alcohol, no turbidity appears at room temperature. Turbidity appears only on heating. If it is a secondary alcohol, turbidity appears in 5 minutes. If it is a tertiary alcohol, turbidity appears immediately.

- c) Methanol and ethanol

Iodoform test: Ethanol when reacted with  $\text{I}_2$  and  $\text{NaOH}$  or  $\text{NaOI}$  gives yellow ppt of iodoform since it has the presence of  $\text{CH}_3\text{-CH(OH)-}$  group.



- **Structure of ethers:**

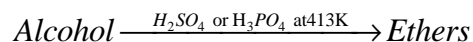


Methoxymethane

(Ether)

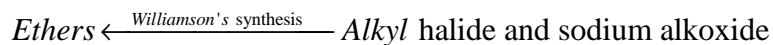
- **Preparation of ethers:**

- a) From alcohols





b) From alkyl halide and sodium alkoxide

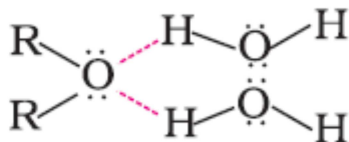


Here, the alkyl halide should be primary and alkoxide should be tertiary. In case of aromatic ether, the aromatic part should be with phenoxide ion.

- **Physical properties of ethers:**

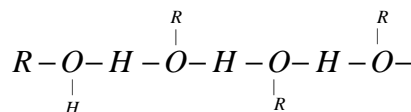
a) Miscibility:

Miscibility of ethers with water resembles those of alcohols of the same molecular mass. This is due to the fact that just like alcohols, oxygen of ether can also form hydrogen bonds with water molecule.



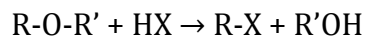
b) Boiling points:

Ethers have much lower boiling points than alcohols. This is due to the presence of hydrogen bonding in alcohols. Hydrogen bonding is absent in ethers.



- **Chemical properties of ethers:**

a) Cleavage of C-O bond in ethers:

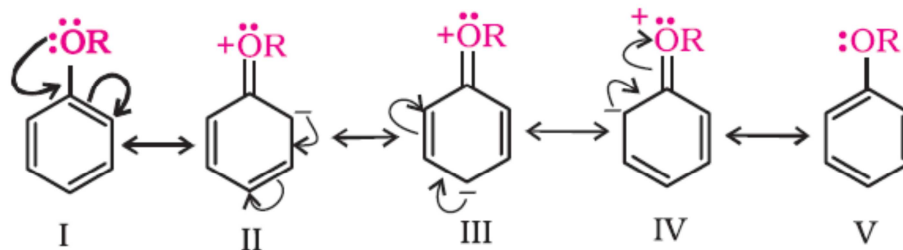


Excess

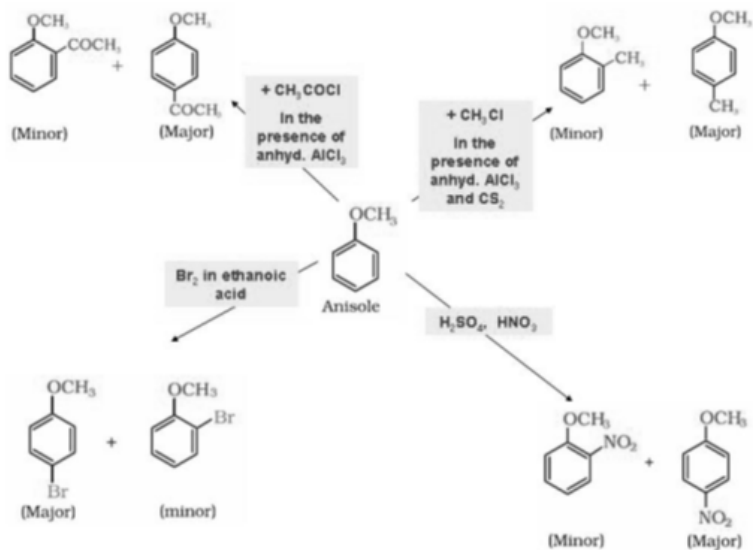
The order of reactivity of hydrogen halides is as follows:  $HI > HBr > HCl$

Alkyl halide formed is always the lower alkyl group. But if a tertiary alkyl group is present, the alkyl halide is always tertiary. In case of phenolic ethers, the cleavage occurs with the formation of phenol and alkyl halide.

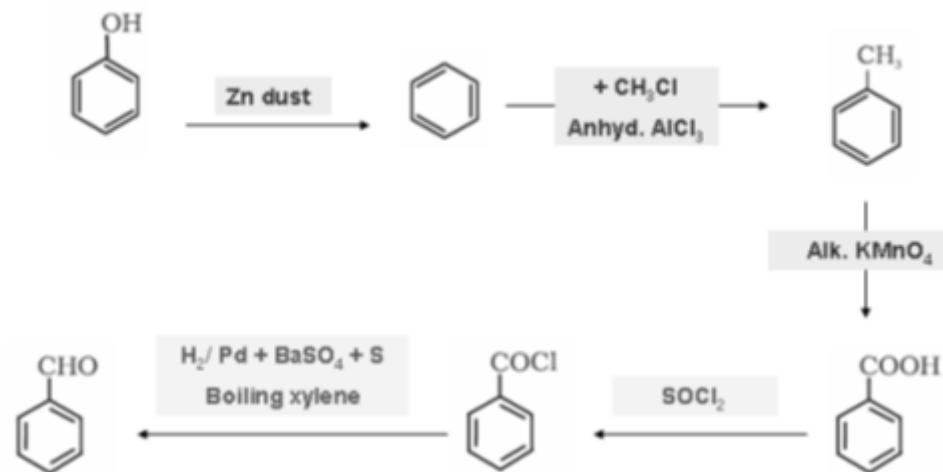
b) Electrophilic substitution reaction in aromatic ethers:



The electrophilic substitution reaction of aromatic ether involves the following reaction:



- Other conversion reactions:
  - Phenol to salicylaldehyde



- Phenol to benzene diazonium chloride

